

The Analytical Chemistry of Uranium

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THE ANALYTICAL CHEMISTRY OF URANIUM



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Theclassics.Us, United States, 2013. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****. This historic book may have numerous typos and missing text. Purchasers can usually download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1903 edition. Excerpt: . Carbon dioxide must also be separately determined. The following analyses are of commercial prepared originally by reducing the chloride with potassium, and this process, or some modification of it, was the only one available until a few years ago. Moissan prepared cast uranium by submitting an intimate mixture of the oxide and sugar charcoal, contained in a coke crucible, to the electric current. He isolated a crystallized carbide (C3U2) whose melting point is much higher than that of platinum, and he prepared samples of the metal containing less than i per cent, of impurities. On this latter metal and other samples prepared by the electrolysis of the double uranium-sodium chloride (UCl4.2NaCl) he determined some of its properties. Ferro-uranium alloys are formed by the last-mentioned process when iron electrodes are used. The metal and some of its alloys have been prepared by various modified forms of the alumino-thermic process: the author, however, found this method to yield poor and very uncertain results. Pure uranium is a white non-magnetic metal, and can be easily filed; it does not scratch glass, is easily carbonized, and may be tempered. Its specific gravity is said to be about 187, but the carbonized metal usually met with has a specific gravity of about 12 only. When in fine powder the metal burns at a comparatively low temperature (170 in pure oxygen), and also very readily combines with nitrogen. It tarnishes on exposure to air, and also decomposes water...

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