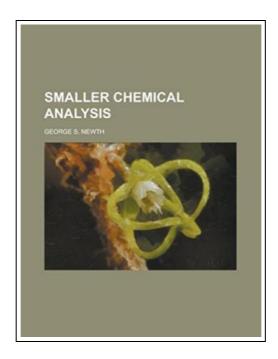
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Reviews

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(Prof. Noah Zemlak DDS)

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RareBooksClub. Paperback. Book Condition: New. This item is printed on demand. Paperback. 46 pages. Dimensions: 9.7in. x 7.4in. x 0.1in.This historic book may have numerous typos and missing text. Purchasers can usually download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1906 edition. Excerpt: . . . CHAPTER VI REACTIONS OF THE METALS OF GROUP Mb Manganese, Mn Dry Reactions. --Manganese compounds, when heated in a borax bead in the oxidising flame, impart to the bead a violet or lilac colour. When heated in the reducing flame, the bead again becomes colourless. A more characteristic reaction is to fuse the manganese compound with KHO, or with Na03 and a little KN03 or KC103 upon a piece of platinum foil; the manganese undergoes oxidation, and a deep green-coloured mass is obtained, consisting of manganates of the alkali metals--Mn02Na2C030 (from KC103or KN03) Na2Mn04C02 Wet Reactions. --Of the common manganous1 salts, the chloride, MnCl2, 4H20, and the sulphate, MnS04, 5H20, are soluble in water. In the crystallised state they have all a pink colour. Manganous salts which are soluble in water do not undergo atmospheric oxidation. NH4HO, KHO, and NaHO produce a white precipitate of manganous hydroxide, Mn(H0)2. Insoluble in excess of the reagent, the precipitated quickly absorbs oxygen, and is converted into hydrated manganic oxide, Mn203, H20, having a brown colour. Freshly precipitated Mn(H0)2, while still white, is soluble in NH4C1; therefore in the presence of NH4C1, manganous hydroxide is not precipitated by NH4H0, and only incompletely by KHO or NaHO. The ammoniacal solution of the double chloride, however, is capable of absorbing oxygen just as the precipitated Mn(H02) 1 The manganic salts are extremely unstable in solution, and do not exist under the ordinary conditions of analysis. does, and the liquid quickly...



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